



$$M_{\text{Cl}} = 9.24\text{g} - 5.78\text{g}$$
$$= 3.46\text{g}$$

$$n_{\text{Sn}} = \frac{m}{M} = \frac{5.78\text{g}}{118.69\text{g/mol}} \quad n_{\text{Cl}} = \frac{3.46\text{g}}{35.45\text{g/mol}}$$

$$0.0486^{(9)} \text{ mol} \quad ; \quad 0.0976^{(10)} \text{ mol}$$

Sn : Cl

$$\frac{0.0486^{(9)}}{0.0486^{(9)}} : \frac{0.0976^{(10)}}{0.0486^{(9)}}$$

1 : 2

