



$$m_{\text{H}_2\text{O}} = 15.82\text{g} - 8.61\text{g}$$
$$= 7.21\text{g}$$

$$n_{\text{NiCl}_2} = \frac{m}{M} = \frac{8.61\text{g}}{129.60\text{g}}$$
$$= 0.0664\text{③ mol}$$

$$n_{\text{H}_2\text{O}} = \frac{7.21\text{g}}{18.02\text{g/mol}}$$
$$= 0.400\text{① mol}$$



$$\frac{0.0664\text{③}}{0.0664\text{③}} : \frac{0.400\text{①}}{0.0664\text{③}}$$

$$1 : 6$$

